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reactions will occur. For each reaction that will occur, complete the chemical equation by writing in the products formed and balancing the final equation. 500C a.  $\text{Al(s)} + \text{CH}_3\text{COOH(aq)}$  500C Thi(s) + 6CH<sub>3</sub>COOH(aq) + 3H<sub>2</sub>(g) 500C b.  $\text{Al(S)} + \text{H}_2\text{O(l)}$  no reaction c.  $\text{cr(s)} + \text{CdC}_{12}\text{(aq)}$   $2\text{Cr(s)} + 3\text{CdC}_{12}\text{(aq)}$   $2\text{CrC}_{13}\text{(aq)} + 3\text{Cd(s)}$  69 d.  $\text{Br}_2\text{(l)} + \text{KC1(aq)}$  no reaction

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